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1 Chemical reactions 1.1 Rate of reaction

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and dependence on temperature. We will once again look at the formation of ammonia ( $\text{NH}_3$ ) from nitrogen and hydrogen (see section Chemical equilibrium of the thermodynamics chapter). This reaction follows the equation:  $\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$  (1)  $\Delta H^\circ = -92 \text{ kJ mol}^{-1}$   $\Delta S^\circ = -192 \text{ J mol}^{-1} \text{ K}^{-1}$  To find the Gibbs

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By sizing a chemical reactor we mean  
we're either determining the reactor volume to  
achieve a given conversion or determine  
the conversion that can be achieved in a



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given reactor type and size. Here we will assume that we will be given  $-r_A = f(X)$  and  $F_{A0}$ . In chapter 3 we show how to find  $-r_A = f(X)$ . Given  $-r_A$  as a function of conversion,  $-r_A = f(X)$ , one can size any type of reactor.

Elements of Chemical Reaction

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Octave Levenspiel was a professor of chemical engineering at Oregon State University. His principal interest was chemical reaction engineering, and he was the author of a major textbook Chemical Reaction Engineering as well as numerous research publications.

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